Worksheet: Oxidation Reduction

1. Indicate whether the following is an oxidation or reduction reac	ction:
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a.
$$O_2(g) + 4e^- \rightarrow 20^{2-}(aq)$$

b. Al(s)
$$\rightarrow$$
 Al³⁺(aq) + 3e⁻

c.
$$2\text{Li}(s) \rightarrow 2\text{Li}^+(aq) + 2e^-$$

d.
$$Cu^{2+}(aq) + 2e^{-} \rightarrow Cu(s)$$

2. In the following reactions, identify which reactant is oxidized and which is reduced:

a.
$$2Mg(s) + O_2(g) \rightarrow 2MgO(s)$$

oxidized ____

reduced _____

b. $2Fe_2O_3(s) \rightarrow 4Fe(s) + 3O_2(g)$

oxidized _____

reduced _____

c. $Cl_2(g) + 2KI(aq) \rightarrow I_2(g) + 2KCl(aq)$

oxidized _____

reduced _____

d. $2PbO(s) \rightarrow 2Pb(s) + O_2(g)$

oxidized ____

reduced