

### Worksheet: Isotopes, Atomic Mass

1. There are four isotopes of sulfur with mass numbers 32, 33, 34, and 36. Write the atomic symbol for each of these isotopes.

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2. Two isotopes of gallium are naturally occurring, with  ${}^{69}_{31}\text{Ga}$  at 60.11% (68.93amu) and  ${}^{71}_{31}\text{Ga}$  at 39.89% (70.92amu). What is the atomic mass of gallium?

3. Given that the atomic mass of Chlorine, Cl, is 35.45amu, solve for the % abundance of the isotope  ${}^{37}_{17}\text{Cl}$ . Chlorine has two isotopes,  ${}^{35}_{17}\text{Cl}$  which weighs 34.97amu at 75.77% and  ${}^{37}_{17}\text{Cl}$  which weighs 36.97amu.