

Worksheet: Periodic Trends

Key

1. List the following in order of increasing atomic radii:

a. K, Na, Rb

Na, K, Rb

b. Be, F, N

F, N, Be

c. N, S, Br

N, S, Br

d. Xe, Ne, Kr

Ne, Kr, Xe

2. List the following in order of increasing ionization energy:

a. Ra, Sr, Ba

Ra, Ba, Sr

b. Bi, P, Sb

Bi, Sb, P

c. F, B, N

B, N, F

d. Si, Ar, S

Si, S, Ar

3. List the following in order of increasing electronegativity:

a. Mg, Ra, Be

Ra, Mg, Be

b. I, F, Cl

I, Cl, F

c. O, Li, C

Li, C, O

d. Fr, S, Sr

Fr, Sr, S

4. Explain why the noble gases do not have electronegativity values.

Already have full octet, do not attract any more electrons.