Worksheet #6 Names

1. Write the molecular equation, ionic equation, and net ionic equation for the reaction between the following solutions. If no precipitate forms, write “no reaction”.

1. KOH(aq) + SrCl2(aq)
2. NH4I(aq) + Pb(ClO4)2(aq)

2. How many moles of NaCl are contained in each volume of aqueous NaCl solution?

1. 2.5L of a 0.25M solution
2. 25 mL of a 2.0M solution
3. 250 mL of a 0.25M solution

3. Classify each substance as a heterogeneous mixture, a solution, or a colloid:

1. Nail polish remover
2. Brass (an alloy of copper and zinc)
3. Milk

4. How many milliliters of 6.0 M NaOH solution would be needed to prepare 525 mL of a 2.5 M solution?

5. Use the solubility rules to predict whether the following ionic compounds are soluble in water:

1. MgCO3 circle one: soluble insoluble
2. PbSO4  circle one: soluble insoluble
3. MgCl2 circle one: soluble insoluble
4. CaCl2 circle one: soluble insoluble

6. How many grams of H2 are formed when 750.mL of 6.00M HCl react?

